Supplementary materials for

Insight into the mechanism of action of marine cytotoxic thiazinoquinones

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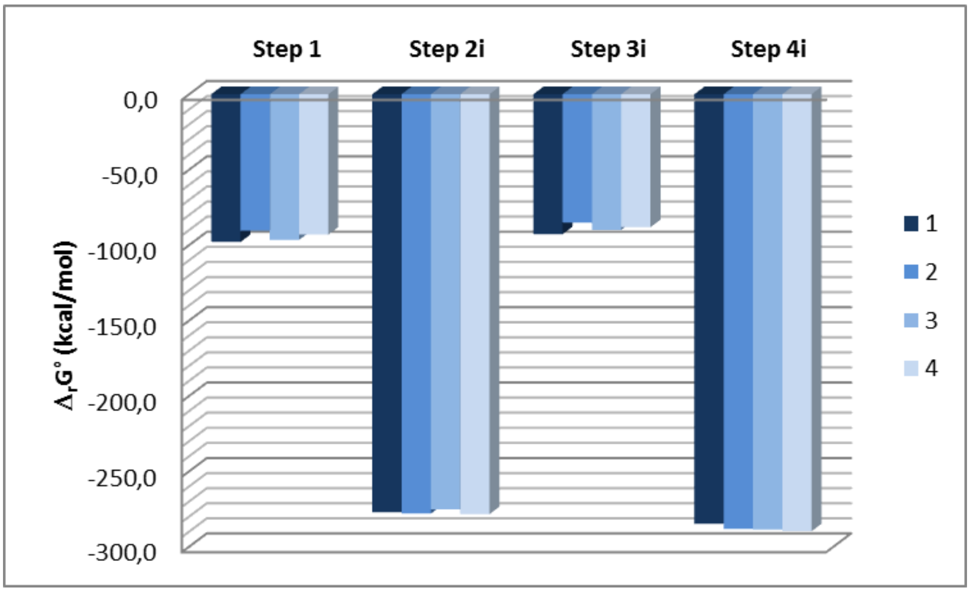

Figure S1 Gibbs free energies of reaction (**Δ_rG°;** kcal/mol) calculated for the proposed reduction pathway considering the formation of the semiquinone species $QH^{\bullet}_{i}$.

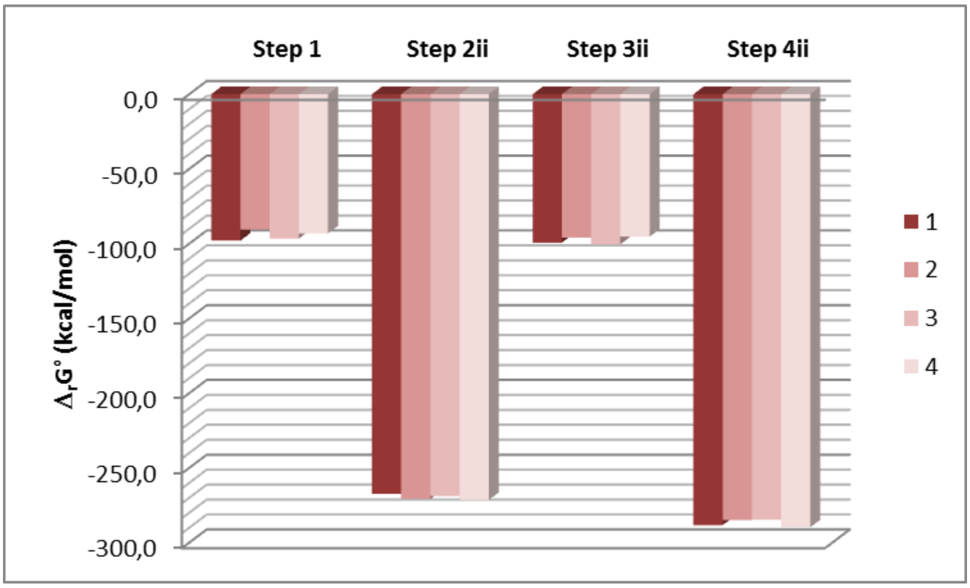

Figure S2. Gibbs free energies of reaction (**Δ_rG°;** kcal/mol) calculated for the proposed reduction pathway considering the formation of the semiquinone species $QH^{\bullet}_{ii}$.

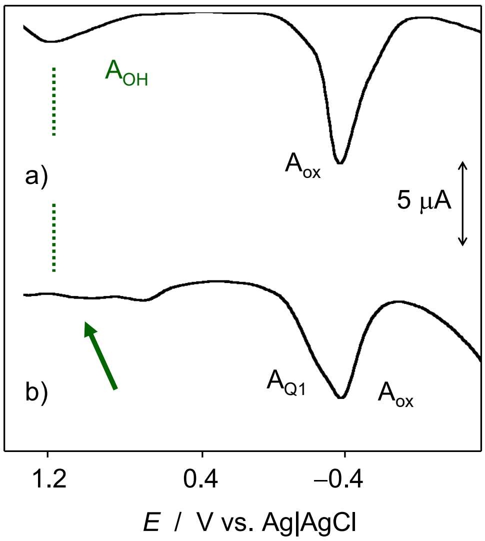

Figure S3. Square wave voltammograms at an air-saturated PBS solution at pH 7.4 at a) unmodified and b) **1**-modified glassy carbon electrode. Potential scan initiated at -0.85 V in the positive direction; potential step increment 4 mV; square wave amplitude 25 mV; frequency 100 Hz. According to Enache et al. [32], the signal A_{OH} appearing at unmodified glassy carbon electrodes corresponds to the one-electron water oxidation yielding the $HO^{\bullet}$ radical. At thiazinoquinone-modified electrodes, this signal is depleted thus suggesting that there is a fast reaction with the Q-compound.